Ν	ar	ne
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PART A – ORBITAL DIAGRAMS & LONGHAND ELECTRON CONFIGURATION

Use the patterns within the periodic table to draw orbital diagrams and write longhand electron configurations for the following atoms.

	Symbol	# e ⁻	Orb	ital Diag	jram and L	onghanc	l Electron	Configura	tion
1.	Mg		1s Elec	2s tron con	2p figuration:	<u>3</u> s	 3p	4s	 3d
2.	Р		 1s Elec	 2s tron con	2p figuration:	<u></u>	 3p	4s	
3.	V		 1s Elec	2s tron con	2p figuration:	<u>3</u> s	3p	4s	 3d
4.	Ge		 1s Elec	2s tron con	2p figuration:	<u>3s</u>	3p	4s	 3d
5.	Kr		 1s Elec	2s tron con	2p figuration:	<u></u>		4s	
6.	Sb		 1s Elec	2s tron con	2p figuration:	<u></u>		4s	

PART B

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IDENTIFY THE ELEMENT DESCRIBED BELOW:

- 1. WHICH ELEMENT CONTAINS A FULL SECOND ENERGY LEVEL?
- 2. WHICH ELEMENT CONTAINS THREE UNPAIRED ELECTRONS IN ITS THIRD ENERGY LEVEL?
- 3. WHICH ELEMENT CONTAINS FIVE ELECTRONS IN ITS 3D ORBITAL?

PART C – RULES OF ELECTRON CONFIGURATIONS

Which of the following "rules" is being violated in each electron configuration below? Explain your answer for each. *Hund's Rule, Pauli Exclusion Principle, Aufbau Principle*

	<u>↑↓</u> 1s	<u>1↓</u> 2s	<u>↑↓</u> 2p				
	<u>1↓</u> 1s		$\frac{\uparrow\downarrow}{2p}\frac{\uparrow\downarrow}{2p}$	 3s	<u>↑↓↑_↑</u> 3p		
9.	<u>1↓</u> 1s	<u>1↓</u> 2s	$\frac{\uparrow\downarrow}{2p}\frac{\uparrow\downarrow}{2p}$		<u>↑↓ ↑↓ ↑_</u> 3p		
	<u>↑↓</u> 1s	<u>↑↓</u> 2s	$\frac{\uparrow\downarrow}{2p}\frac{\uparrow\downarrow}{2p}$	<u>↑↓</u> 3s	<u>↑↓ ↑↓ ↑↓</u> 3p	$\frac{\uparrow\downarrow\uparrow\uparrow\downarrow\uparrow\downarrow\uparrow\downarrow\uparrow\downarrow\uparrow\downarrow}{3d}$	